

Table 4. Standard redox potentials of different electron donors of the photosynthetic light reaction.^a

Electron donor	E_o^- [mV]
$1/2O_2/H_2O$	+820
$Fe(OH)_3 + HCO_3^-/FeCO_3$	+200
Fumarate/Succinate	+33
HSO_3^-/S^0	-38
SO_4^{2-}/S^0	-200
SO_4^{2-}/HS^-	-218
$Fe(OH)_3/Fe^{2-}$	-236
S^0/HS^-	-278
$HCO_3^-/acetate$	-350
$S_2O_3^{2-}/HS^- + HSO_3^-$	-402
$H^+/1/2H_2$	-414
Electron acceptor	E_o^- [mV]
$CO_2/<CH_2O>$	-434

^aTaken from Brune, 1989; Widdel et al., 1993; Thauer et al., 1977; Zehnder and Stumm, 1988.