

Table A2 Midpoint redox potentials for reactions of interest in photosynthesis

Redox reaction	E_m' (V) ^a
$\text{NADP}^+ + \text{H}^+ + 2\text{e}^- \rightleftharpoons \text{NADPH}$	-0.324
$\text{O}_2 + 2\text{H}^+ + 4\text{e}^- \rightleftharpoons 2\text{H}_2\text{O}$	+0.816
$\text{P700}^+ + \text{e}^- \rightleftharpoons \text{P700}$	+0.49
$\text{P870}^+ + \text{e}^- \rightleftharpoons \text{P870}$	+0.45
$\text{P680}^+ + \text{e}^- \rightleftharpoons \text{P680}$	-1.1
$2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{H}_2 \text{ (g)}$	-0.414
$\text{UQ} + 2\text{e}^- + 2\text{H}^+ \rightleftharpoons \text{UQH}_2$	+0.060
$\text{Chl}^+ + \text{e}^- \rightleftharpoons \text{Chl}$	+0.78
$\text{BChl}^+ + \text{e}^- \rightleftharpoons \text{BChl}$	+0.64

^a All values refer to the standard state of 298 K and pH 7.